

>> Here's two kinds of replacement reactions, a single replacement and a double replacement. I want to go over these so that you can understand the difference between the two because you're going to have to identify which is single and which is double. The top one is a single replacement. With a single replacement, you start with an element and a compound. Ok? What you end up with is the element and one of the components of the compound changing places with each other. So when you look at your results over here, your products, you again have an element and a compound, but now they've exchanged places. What had once been an element is now part of the compound. What had been the compound is now an element. Ok. So you're starting with an element and a compound. Now with double replacement, you're seeing a switching of places again, but look what you're starting with. A compound and a compound. You don't have an element. Remember, an element is something that's just by itself. Iron here is just by itself. Copper is just by itself. Anytime you have two different elements together, it makes a compound. So you've got a compound and a compound. And the metal of each of the compounds are switching places with each other to now have different partners. So this is your double placement. It only has compounds switching, with some switching going on. Single replacement has an element switching with a compound.